SNC2D Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Chemistry Unit – What you need to know**

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| We covered material in the following sections: | Try these review questions in the textbook: |
| -4.1, 4.2, 4.3  -5.1, 5.2, 5.3  -6.1, 6.2, 6.3 | Chapter 4: p.174-175  Chapter 5: p.214-215  Chapter 6: p.252-253  Unit 2: p.258-260 |

*You should be able to:*

-draw Bohr-Rutherford & Lewis dot diagrams of common ionic and molecular compounds

[review from Grade 9]

-name, determine chemical formulae for & composition of:

-binary ionic compounds & ternary compounds

(including multivalent metals and polyatomic ions)

-binary molecular compounds

-binary acids, oxoacids, and bases

-use data collected from a chemical reaction to support the law of conservation of mass

-explain the rationale for balancing chemical equations

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simple chemical reactions

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-identify types of reactions (synthesis, decomposition, single displacement, double displacement) and represent them using chemical equations

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-describe the reactants & products of acid-base neutralization

-describe how the pH scale is used to classify solutions as acidic, basic, or neutral

-name common acid-base indicators & other methods of determining the pH of a solution

-determine the pH of a solution using colour changes in red/blue litmus paper, bromothymol

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